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Sections 4 1 4 2 General Properties of Aqueous Solutions and Precipitation Reactions

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Note: Not all of the reactions will occur. For those that do not, write no reaction.

1. A piece of copper is dropped into a container of water. No reaction
2. Liquid bromine is added to a container of sodium iodide crystals. $\text{Br (l)} + 2\text{NaI(s)} \rightarrow 2\text{NaBr(s)} + \text{I}_2\text{(s)}$
3. An aluminum strip is immersed in a solution of silver nitrate.

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tin(II) carbonate sodium hydrogen carbonate
NaHCO₃ 1. 2. 3. 4. 5. 6. 8. 9. 10. 11.
vanadium(V) oxide H₂O dihydrogen monoxide
(NH₄)₂CO₃ ammonium oxalate 4 224 polonium(VI)
thiocyanate Po(SCN)₆ tetraphosphorus
decaoxide P₄O₁₀ 12. 13. 15. 16. 17. 18.
19. manganese(VII) oxide copper(II)
dihydrogen phosphate

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determination of products of chemical reactions (the Equations section) five of these eight equations to answer (Note: Additional responses will not be graded or counted If all [PDF] Ultimate Chemical Equations Handbook Answer Key

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